

University of Mumbai
Examination 2020- Inter Cluster

Program: BE Instrumentation Engineering

Curriculum Scheme: Revised 2016

Examination: Second Year Semester IV

Course Code and Course Name: ISC404 and Analytical Instrumentation

Time: 1 hour

Max. Marks: 50

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Note to the students:- All Questions are compulsory and carry equal marks .

Q1.	Spectroscopy deals with interaction of electromagnetic radiation with matter. What is the speed of this radiation in vacuum in m/s?
Option A:	6×10^8
Option B:	5×10^8
Option C:	7×10^8
Option D:	3×10^8
Q2.	Which type of Quantum Transition takes place in Ultra Violet and Visible spectroscopy?
Option A:	Rotation of molecules
Option B:	Nuclear
Option C:	Bonding electrons
Option D:	Spin of nuclei in magnetic field
Q3.	Which of the following is not a property or parameter of electromagnetic radiation?
Option A:	Wavelength
Option B:	Voltage
Option C:	Wave number
Option D:	Amplitude
Q4.	Which of the following is not a type of Spectroscopy?
Option A:	Gamma ray

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Option B:	X ray
Option C:	Nuclear magnetic resonance
Option D:	Sound
Q5.	Which of the following is false about wavelengths of electromagnetic radiation?
Option A:	Radiation with short wavelengths have high energies
Option B:	Energy does not depend on wavelength
Option C:	Radiation with long wavelengths have low energies
Option D:	Energy depends on wavelength
Q6.	Which of the following detectors does not require a battery and is also known as barrier layer cell?
Option A:	Photomultiplier tube
Option B:	Photovoltaic cell
Option C:	Photoemissive tubes
Option D:	Photo reflector
Q7.	Which of the following detectors is used to detect light intensities which are very weak?
Option A:	Photomultiplier tube
Option B:	Photovoltaic cell
Option C:	Photoemissive tubes
Option D:	Photo reflector
Q8.	How does continuous wedge filter differ from normal interference filter used in absorption spectroscopy?
Option A:	It permits continuous selection of different wavelength
Option B:	It allows narrow band of wavelengths to pass
Option C:	It has two semi-transparent layers of silver
Option D:	Space layer is made of a substance having low refractive index
Q9.	Which of the following statements is false about single beam absorption instruments?
Option A:	Tungsten bulb is used as source

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Option B:	Beamsplitter is used to get parallel beam
Option C:	Test tube is used as sample holder
Option D:	Photovoltaic cell as detector
Q10.	Which of the following statement is false about double beam absorption instruments?
Option A:	It is similar to single beam instruments except two beams are present
Option B:	Tungsten bulb is used as source
Option C:	Reference beam must have higher intensity than sample beam
Option D:	Both the beams after they pass through respective samples are compared
Q11.	Which of the following is not an application of colorimeter?
Option A:	Paints
Option B:	Inks
Option C:	Cosmetics
Option D:	Composition detection
Q12.	In photometers, the readings of the specimen are initially obtained in the form of which of the following parameters?
Option A:	Transmittance
Option B:	Absorption
Option C:	Wavelengths
Option D:	Volume
Q13.	Which of the following is the purpose of balance indicator in double beam photometer or colorimeter?
Option A:	Selects particular wavelength
Option B:	Splits the wavelength selected into two equal beams
Option C:	Detects and indicates the amount of light falling on it
Option D:	Indicates the difference between the output of two photometers
Q14.	Which of the following is not true about Fourier Transform Infrared (FTIR) spectrometer?

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Option A:	It is of non-dispersive type
Option B:	It is useful where repetitive analysis is required
Option C:	Size has been reduced over the years
Option D:	Size has increased over the years
Q15.	Only pyroelectric transducer or pyroelectric crystals are used as detectors in Fourier Transform Infrared Spectrophotometer (FTIR). What is the main reason for other types of thermal detectors are not being used in FTIR spectrophotometer?
Option A:	Less accuracy
Option B:	Slower response
Option C:	Less precision
Option D:	Less sensitivity
Q16.	Chromatography is a physical method that is used to separate and analyse _____
Option A:	Simple mixtures
Option B:	Complex mixtures
Option C:	Viscous mixtures
Option D:	Metals
Q17.	In which type of chromatography, the stationary phase held in a narrow tube and the mobile phase is forced through it under pressure?
Option A:	Column chromatography
Option B:	Planar chromatography
Option C:	Liquid chromatography
Option D:	Gas chromatography
Q18.	In chromatography, the stationary phase can be _____ supported on a solid.
Option A:	Solid or liquid
Option B:	Liquid or gas
Option C:	Solid only
Option D:	Liquid only
Q19.	In chromatography, which of the following can the mobile phase be made of?
Option A:	Solid or liquid

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Option B:	Liquid or gas
Option C:	Gas only
Option D:	Liquid only
Q20.	Nitrogen oxide cannot be directly analyzed using UV and Visible analyzers due to which of the following reasons?
Option A:	Less accuracy
Option B:	Very low range
Option C:	It leads to contamination of the sample
Option D:	It is transparent in UV visible regions
Q21.	How is NO converted to NO ₂ for analysis in UV and Visible analyzers?
Option A:	Treating sample gas with pressurized oxygen
Option B:	Treating sample gas with ozone
Option C:	Treating sample gas with oxygen at low pressure
Option D:	Treating sample gas with water at high pressure
Q22.	Which of the following is the full scale range of Infrared spectroscopy method used for measurement of carbon monoxide?
Option A:	0-2 ppm
Option B:	0-500 ppm
Option C:	0-1 ppm
Option D:	0-50 ppm
Q23.	Which of the following is the full scale range of Gas chromatography method used for measurement of carbon monoxide?
Option A:	0-200 ppm
Option B:	0-500 ppm
Option C:	0-1 ppm
Option D:	0-50 ppm
Q24.	Mass spectrometers are used to determine which of the following?
Option A:	Composition in sample
Option B:	Concentration of elements in sample
Option C:	Relative mass of atoms
Option D:	Properties of sample

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Q25.	Mass spectrometer separates ions on the basis of which of the following?
Option A:	Mass
Option B:	Charge
Option C:	Molecular weight
Option D:	Mass to charge ratio